

## CLASSROOM EXPECTATIONS

### General Course Objectives:

Over the course of the semester, the chemistry student will understand the structure and function of the basic principles of matter. Through research, experimentation, and exploration, the chemistry student will gain understanding of the properties of atoms and atomic interaction.

3 units first semester: periodic table, bonding, stoichiometry

3 units second semester: thermodynamics, rates & equilibrium, nuclear chem

### Materials:

- folder
- notebook (1 subject)
- pens/pencils
- loose-leaf paper
- **CALCULATOR**
- misc lab supplies
- notecards

### Grading:

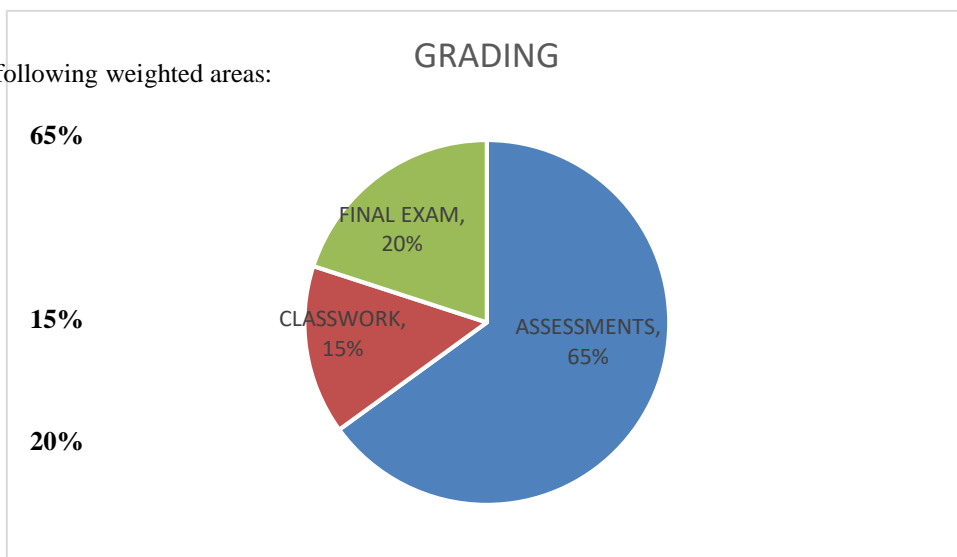
Students will be graded on their mastery of curriculum, skills, and the ability to analyze data. Successful students maintain focus during class and spend time outside of class reviewing concepts and practicing skills.

Grading is on a straight scale.

90 - 100 % = **A**    80 - 89.9% = **B**    70 - 79.9% = **C**    60 - 69.9% = **D**    below 60% = **F**

Grades will be based on point accumulation in the following weighted areas:

- **ASSESSMENT**  
TESTS (1 PER UNIT)  
QUESTS (1 PER UNIT)  
TYPED LAB REPORTS (2 PER UNIT)  
PROJECTS
- **CLASSWORK**  
OPEN NOTE QUIZZES  
MINI LABS & ACTIVITIES
- **FINAL EXAM**



### Extra Credit:

Extra credit can be earned on lab reports. +1 for further research on lab topic, +1 for related picture  
PARENT NEWSLETTER PROJECT (1 PER SEMESTER)

### Academic Integrity:

Refer to PNHS online student handbook and dishonesty policy

-3 levels of indiscretions

-consequences range from zero on assignment to out of school suspension

Open note quizzes – Level I

Tests, Quests & Labs – Level II

Occasionally lab data will be identical to lab partners, but analysis of lab data must be your own work.





## CLASSROOM BEHAVIOR

1. Be prepared. Bring **ALL** necessary materials to class including textbook, pencils, pens, science notebook, folder, **calculator**, etc.
2. Please be seated when bell rings.
3. Please raise your hand when you wish to speak.
4. Absolutely **NO** food or drinks will be permitted in class.
5. **NO** cell phone use during class.
6. Treat all students and staff with respect.
7. During tests, cell phones are to be turned off and stored in backpacks at front of class room.
8. **All handbook rules will be enforced!**



Violation of the expected behaviors will be dealt with in a four step process:

- 1) verbal warning
- 2) parent contact and/or student-teacher conference
- 3) parent contact and 30 minute detention (beaker duty)
- 4) parent contact and dean referral

\*\*\*\*\*

## CLASSROOM PROCEDURES

### Beginning of Class:

All students must be seated when the bell rings. A warm up activity will be on the board. Be sure to check weekly calendar on board for important due dates.

### Tardies:

Tardy students will not be allowed into classroom without a pass from dean.

### Absences:

If you are absent, please refer to the honors chem bin which has a calendar summarizing class and extra handouts. Please come to class early to get absent work.

-A weekly schedule will also posted on classroom whiteboard for student reference.

-Students have one day to make up work for each day of an excused absence per student handbook policy.

If you are absent for an open note quiz, you do not have to make it up. If you are absent for a lab, please get sample data from instructor or teacher pages. Absence students are still required to turn in lab report. If you are absent from a test or quest, please schedule make up with teacher.

### Turning in work:

Please put work to be turned in (including late work) in the turn it in shelves.

### Hall passes:

Hall passes will be given to a student who needs to leave the classroom for emergencies only. Students must complete the sign out log.

**NO LOCKER PASSES!**

### Lab Expectations:

Lab **safety** is the number one **priority** of any science teacher. For the safety of all students, it is imperative that all lab directions are followed and the lab safety contract is upheld. Any student who acts in a manner considered dangerous will be removed from the lab and given a **detention or referral**.

### Late Policy:

Late labs and projects will be penalized 10% for each day late. **No** late work (including labs) will be accepted after the unit test. Unit study guides are due the day before the test. Study guides won't be accepted for partial/late credit.

## FREQUENTLY ASKED QUESTIONS

1. IS THE HONORS CHEMISTRY TEXTBOOK AVAILABLE ONLINE? No
2. WHAT RESOURCES ARE AVAILABLE?  
Read the textbook or research on the internet.  
The teacher pages will have notes & practice problems with solutions.  
Mrs. McBride is available after school or during study hall for help.  
Peer tutors are available in the media center during lunch study halls.
3. DO YOU GIVE TEST RETAKES OR TEST CORRECTION? No
4. WHAT CAN I DO TO EARN AN "A"?  
"A" students consistently score above a 90% on tests, quests, lab reports and the final exam  
"A" students review material outside of class regularly
5. WHAT CAN I DO TO BRING MY GRADE UP?  
Review notes and practice worksheets outside of class regularly  
Meet with Mrs. McBride for extra help after school  
Ask questions in class  
Complete the extra credit newsletter  
Add extra details or appropriate pictures to lab reports.
6. IS THERE DAILY HOMEWORK?  
No, no work is assigned for points. However, students are expected to review material & skills outside of class.  
Additionally, some students may need to complete lab reports outside of class.
7. HOW MUCH TIME IS GIVEN FOR WRITING LAB REPORTS.  
Usually students are provided 1 day of class time to type lab report. Lab reports are due 1 week from day of data collection.

## 2019-2020 Units of Study and Main Topics for HONORS CHEMISTRY

### UNIT 0 : MEASUREMENT

Lab safety, Metric system, significant figures, scientific notation

### UNIT 1: PERIODIC TABLE

Atomic theory, electron configuration, periodic table trends, periodic table families

### UNIT 2: CHEMICAL BONDS

Lewis dot structures, chemical bonds, naming compounds, writing chemical formulas, intermolecular forces

### UNIT 3: CHEMICAL REACTIONS AND STOICHIOMETRY

Molar mass, mole and gram conversions, empirical formulas, molecular formulas, balancing chemical equations, types of chemical reactions, percent composition, theoretical yield, percentage yield

### UNIT 4: THERMODYNAMICS

Specific heat, comparing endothermic and exothermic reactions, potential energy diagrams, heat of reaction, molar heat of fusion, heating curves, phases of matter

### UNIT 5: CHEMICAL EQUILIBRIUM

Reaction rates, potential energy diagrams, chemical equilibrium, molarity, molality, colligative properties, titrations, solutions, solubility curves

### UNIT 6: NUCLEAR CHEMISTRY

Alpha, beta, and gamma radiation, balancing nuclear equations, comparing fission to fusion, half life, star cycle, and effects of radiation

Please retain a copy of this document for your records. Document is also available on teacher pages.

Please email Mrs. McBride with questions or concerns at: [nmcbride@psd202.org](mailto:nmcbride@psd202.org)

Or the math & science divisional chair Mrs. Brest at [gbrest@psd202.org](mailto:gbrest@psd202.org)